

Abstract Submitted
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Protons Are Spinning On regular Orbits And Neutrons Have Different Binding Energies Inside Atoms NIAZI ELFIKKY, Saudi Electricity Company — Preface. With the centennial of the discovery of the atomic nucleus by the Chemist, Professor Ernest Rutherford (1871-1937), it is worthwhile to memorize such occasion. For the structure of the atom where he had postulated that atoms have their positive charge concentrated in a very small nucleus. It was a great breakthrough when he introduced his imagination for the structure of the atom. Furthermore, he proposed a dynamic planetary model in which the nucleus plays the role of the sun and electrons correspond to the individual planets of the solar system although he fixed the atom's nucleus of the solar system, (He resembled the nucleus of the atom with the sun although he had dropped an essential fact that the sun is also rotating in its specific orbit surrounded by the outer orbit for the rotating planets like the whirling electrons. For studying the stable atom of the steady state elements like hydrogen $1H1$, the particles of the $1H1$ atoms are simple just one proton $1p1$ and one electron $0e1$, both the proton and electron have equal electric charge ($q= 1.6 \times 10^{-19}$ Coulomb) but in different sign (positive charge for proton and negative for electron), also they have different masses ($m_p= 1.67 \times 10^{-27} \text{Kg}$, $m_e= 9.11 \times 10^{-31} \text{kg}$), proton is heavier than electron by 1.834×10^3 . What does it mean?

Niazi Elfikky
Saudi Electricity Company

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